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## Convenient synthesis of EDOT-based dyes by C–H activation and their application as dyes in dye-sensitized solar cells†

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Precursors to three new 3,4-ethylenedioxythiophene (EDOT) incorporating dyes have been synthesised via a one-pot C–H activation route using *N,N*-dimethylaniline as a donor group. We have extended this methodology to provide a convenient one-pot route to dye **EDOT-Ph**. The electrochemical and optical properties of the new dyes have been correlated with IV and EQE data for 1 cm<sup>2</sup> dye-sensitized solar cell (DSSC) devices prepared using these dyes. The device data show that dye performance is strongly affected by the amount of chenodeoxycholic acid (CDCA) co-sorbent used. The best performance is for **EDOT-Ph** ( $\eta$  = 4.0%) at 10 mM CDCA compared to ( $\eta$  = 6.0% and  $\eta$  = 5.8%) for **N719** and **D205** control cells.

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## Introduction

In recent years, dye-sensitized solar cell (DSSC) devices fabricated using metal-free organic dyes have begun to overtake metal-based dyes in terms of device performance. For instance, there have been several reports of organic dyes with power conversion efficiencies ( $\eta$ ) at 1 Sun > 12% (ref. 1–4) with the best non-verified DSSC device efficiency to date ( $\eta$  = 14.7%) being for a device co-sensitized with the alkoxysilyl anchor dye **ADEKA-1** and a cyanoacrylate-functionalised dye **LEG4**.<sup>5</sup> The interest in metal-based dyes stems from the initial report by O'Regan and Grätzel using a Ru-bipyridyl complex (**N3**) to sensitize TiO<sub>2</sub> nanoparticles giving  $\eta$  > 7%.<sup>6</sup> The closely related Ru-bipy dye (**N719**)<sup>7</sup> increased efficiency whilst the alkyl-modified **Z907** improved dye stability to water ingress.<sup>8</sup> To date, the best Ru-bipy DSSC dye (**C106**) gives  $\eta$  = 11.4%.<sup>9</sup> However, whilst Ru-bipy DSSC dyes are effective, their spectral response is generally restricted to <650 nm (with the exception of the related terpyridyl "Black dye"<sup>10</sup>), molar extinction coefficients ( $\epsilon$ ) are much lower than organic dyes and potential structural modifications are limited to the periphery of a relatively small

number of metal-chelating ligands suitable to make stable complexes.

By comparison, organic DSSC dyes can be tuned to absorb across the visible spectrum from yellow<sup>11</sup> to red<sup>12,13</sup> to blue,<sup>14,15</sup> they possess much higher  $\epsilon$  (e.g. 14 000 M<sup>−1</sup> cm<sup>−1</sup> for **N719** (ref. 16) vs. 330 000 M<sup>−1</sup> cm<sup>−1</sup> for **SQ2** squaraine<sup>15</sup>) and they offer an almost infinite range of potential structural modifications although generally a donor-spacer-acceptor arrangement of groups in the dye is most widely utilised.<sup>17</sup> As such, organic DSSC dyes are best grouped into dye families based on the chromophore; e.g. triphenylamines,<sup>18</sup> indolines,<sup>19,20</sup> coumarins<sup>21</sup> as examples of dye which harvest light <650 nm and squaraines,<sup>14,15</sup> phthalocyanines<sup>22,23</sup> or cyanine<sup>24</sup> dyes which absorb  $\lambda$  > 650 nm. To extend light harvesting across the visible spectrum, devices can be co-sensitized with two<sup>25</sup> or more dyes.<sup>11,26</sup> Another approach is to develop panchromatic dyes such as panchromatic squaraines.<sup>27</sup> However, these dyes tend to contain many different groups which complicates synthesis and purification<sup>28</sup> and ultimately increases dye cost which is an issue for device scaling and commercialisation.

The majority of organic dyes for DSSCs are synthesised using standard Pd-catalysed aryl–aryl coupling reactions such as Suzuki<sup>29</sup> and Stille<sup>30</sup> coupling methodologies. However, the application of direct C–H activation is an attractive strategy to form new C–C bonds as it offers a largely unlimited range of commercially available substrates, and removes the necessity of preparing organoboron or toxic organotin reagents as is the case for Suzuki and Stille coupling reactions, respectively.<sup>31</sup>

This paper describes the synthesis of EDOT dyes using one-pot Pd-catalysed C–H activation protocols to afford unsymmetrical precursors to three new 3,4-ethylenedioxythiophene (EDOT)-bridged dyes **EDOT-Ph**, **EDOT-Fu** and **EDOT-Th** (Scheme 1).

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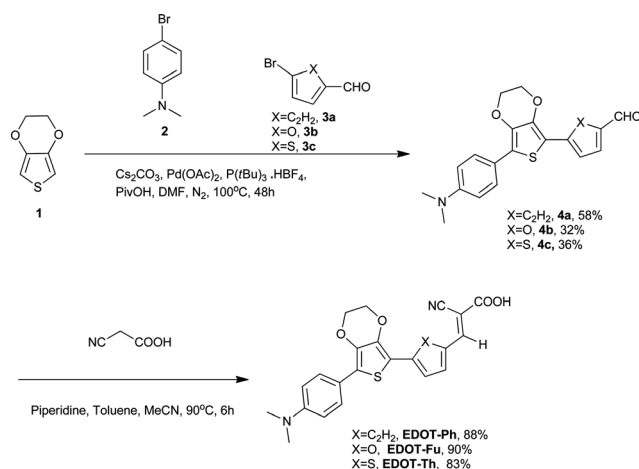
† Electronic supplementary information (ESI) available: Synthetic method development, square wave voltammetry, device characterization and testing, UV-vis spectra of thin films of the dyes and solution NMR spectra. See DOI: 10.1039/c6ta04483b

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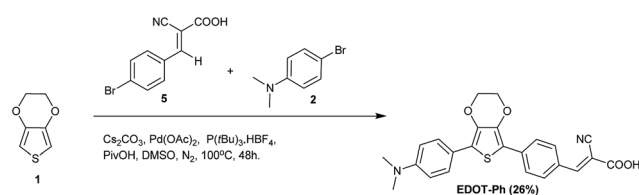


Moreover, we have extended this methodology to prepare the best performing dye (**EDOT-Ph**) in a one-pot reaction of EDOT and commercially available donor and acceptor residues (Scheme 2).<sup>32</sup> Although EDOT has been shown to be a good substrate for Pd-mediated C–H activation reactions<sup>33</sup> and has been used to prepare unsymmetrical push–pull dyes by the sequential addition of donor and acceptor units,<sup>34</sup> our one-pot approach<sup>35</sup> described here significantly reduces the synthesis time and associated consumables costs relating to the isolation of intermediates, allowing the rapid and convenient synthesis of **EDOT-Ph** in reasonable yield.

EDOT has been used widely in DSSC dyes as a conjugated bridge unit between the donor and acceptor residues due to its ability to facilitate co-planarity and improve light harvesting.<sup>36</sup> Previous reports of EDOT-based DSSC dyes have typically relied on 2 or 3-step syntheses using a *N,N*-dimethylaniline donor bonded to EDOT with either a thiophene-acetylene linker ( $\eta = 4.5\%$ )<sup>28</sup> or with a triphenylamine donor bonded to EDOT with a thiophene linker ( $\eta = 6.0\%$ )<sup>36</sup> or a fluorinated phenyl linker ( $\eta = 8.2\%$ ).<sup>37</sup> Furthermore, the 8.2% efficiency reported<sup>37</sup> was achieved by modification of the acceptor moiety of the sensitizer. The new dyes and protocols described here give comparable device performance to related sensitizers,<sup>36,37</sup> but have been synthesized using a convenient one-pot C–H activation methodology using straightforward protocols and inexpensive purification. In addition, these *N,N*-dimethylaniline-containing EDOT dyes give higher device efficiencies than the aforementioned sensitizers when compared to **N719** in a side-by-side comparative study. Furthermore, the devices used in this study are relatively large *i.e.*  $1\text{ cm}^2$  *cf.*  $0.25\text{ cm}^2$ .<sup>28</sup>



**Scheme 1** One-pot synthesis of **4a–c** via C–H activation of EDOT and their subsequent conversion to dyes **EDOT-Ph**, **EDOT-Fu**, **EDOT-Th**.



**Scheme 2** One-pot synthesis of **EDOT-Ph**.

## Results and discussion

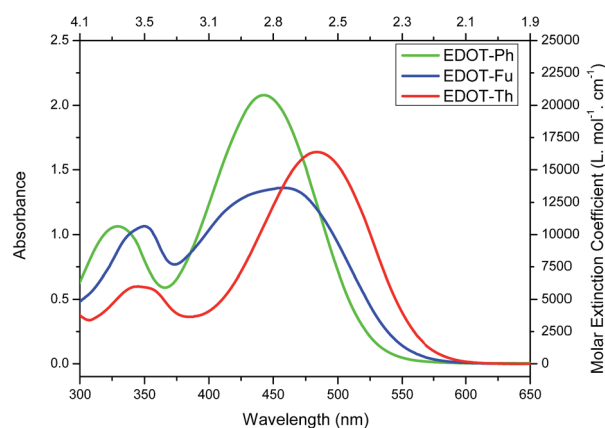
### Synthesis

One-pot C–H activation was used to prepare a series of dye precursors for DSSC devices using *N,N*-dimethylaniline as a donor, EDOT as the  $\pi$ -bridge and three different aromatic/heteroaromatic aldehydes (Scheme 1). A range of conditions were investigated to maximize the yields of these reactions, and a summary is reported in the ESI.† The best conditions gave compounds **4a–c** in 32–58% yield which offers convenient and rapid access to the dye precursors in a one-pot procedure. The aldehyde moiety was then reacted with cyanoacetic acid to provide three dyes **EDOT-Ph**, **EDOT-Fu** and **EDOT-Th** in high yields (Scheme 1). Furthermore, we were able to further exploit this methodology to prepare **EDOT-Ph** in a one-pot protocol in a respectable yield (Scheme 2). It is noteworthy that it was not necessary to protect the carboxylic acid unit of compound **5** which thus avoided an additional de-protection step, which is common for many DSSC dyes of this type. Whilst protection of carboxylic acids is synthetically simple to achieve, the addition and removal requires additional synthetic steps and complicates routes to high-purity dyes which perform best in devices.

### Optical and electrochemical properties

All three dyes were then characterised by UV-vis spectroscopy (Fig. 1) and square-wave voltammetry (Fig. S1†) and the data are summarised in Table 1. The three dyes show strong absorbance across the visible range with maxima of absorption between 444–483 nm as well as good molar extinction coefficients ranging from  $13\,630$ – $20\,790\text{ L mol}^{-1}\text{ cm}^{-1}$ . The UV-vis spectra show a red shift in the onset of absorbance on going from **EDOT-Ph** to **EDOT-Fu** and **EDOT-Th** suggesting a narrowing of the optical band gap ( $E_{\text{opt}}$ ) from 2.38 eV (**EDOT-Ph**) to 2.20 eV (**EDOT-Th**).

We have utilised square-wave voltammetry to determine the ionization potential (IP), electron affinity (EA) and the fundamental gap ( $E_{\text{fund}}$ ) where possible. The IP of the three dyes were relatively similar with a difference of *ca.* 50 mV. The reduction waves for **EDOT-Ph** were better resolved than the other dyes



**Fig. 1** UV-vis absorbance spectra of dyes **EDOT-Ph**, **EDOT-Fu** and **EDOT-Th** ( $\text{THF } 1 \times 10^{-4}\text{ M}$ ).



Table 1 Electrochemical and optical data (V vs. Fc/Fc<sup>+</sup>)<sup>a</sup>

Dye	EDOT-Ph	EDOT-Fu	EDOT-Th
<b>Electrochemical properties</b>			
$E_{\text{ox}}$ (V)	0.48	0.51	0.46
$E_{\text{red}}$ (V)	−1.90	×	×
IP (eV)	−5.28	−5.31	−5.26
EA (eV)	−2.90	×	×
$E_{\text{fund}}$ (eV)	2.38	×	×
<b>Optical properties</b>			
$\lambda_{\text{max}}$ (nm)	444 ( $\epsilon = 20\,790$ )	457 ( $\epsilon = 13\,620$ )	483 ( $\epsilon = 16\,380$ )
$\lambda_{\text{onset}}$ (nm)	522	548	564
$E_{\text{opt}}$ (eV)	2.38	2.26	2.20

<sup>a</sup> Extinction coefficient is given in  $\text{L mol}^{-1} \text{cm}^{-1}$  (from  $A = \epsilon cl$ ). × = not calculated.  $E_{\text{opt}}$  is obtained from  $E = 1240/\lambda_{\text{max}}$ .

allowing the estimation of its EA (−2.90 eV). This gives an  $E_{\text{fund}}$  of 2.38 eV for this dye which is consistent with the optically determined gap. The relatively unchanged IP and the optically measured narrowing of the energy gap suggest that the change in substituents chiefly affects the LUMO/EA of the three dyes.

### Theoretical calculations

DFT calculations indicate that all dyes have essentially identical HOMO and LUMO energies (*ca.* −4.9 eV, −2.4 eV, respectively). The isosurface maps indicate reasonable overlap for the HOMO/LUMO, with the former being mainly located over the donor *N,N*-dimethylaniline and EDOT units, whereas the LUMO is largely localised over the cyanoacrylate and its adjacent aromatic linker unit (Fig. 2). The dyes are reasonably planar with **EDOT-Ph** providing the lowest calculated dihedral angle between the *N,N*-dimethylaniline and EDOT units (16.7° *versus*

~17.1° for the other dyes), whilst **EDOT-Th** has the smallest dihedral angle between the EDOT and the aromatic linker to the cyanoacrylate residue (0.5° *versus* 11.1° for **EDOT-Ph** and 4.8° for **EDOT-Fu**). Interestingly, the dipole moments show some variation with **EDOT-Th** having the largest dipole moment of 11.6 Debyes and **EDOT-Ph** and **EDOT-Fu** possessing slightly lower values of 11.3 and 11.2 Debyes, respectively.

### Device testing

Table 2 and Fig. 3 show *I*–*V* data for the devices prepared using the EDOT dyes with and without 10 mM chenodeoxycholic acid (CDCA) and/or 0.05 mM LiI. The data show that working DSSC devices are produced from all variables with comparable performance from **EDOT-Ph** and **EDOT-Fu** without any CDCA or LiI ( $\eta = 1.0\%$  and  $1.2\%$ , respectively). However, adding CDCA has a profound influence on device efficiencies for all the dyes improving  $\eta$  by 2–5 times to 3.9%, 2.4% and 2.7% for **EDOT-Ph**, **EDOT-Fu** and **EDOT-Th**, respectively.

We have observed the positive effect of CDCA on device performance previously but not to this extent.<sup>11–13,25</sup> For all the dyes, the improvements reflected increased  $V_{\text{oc}}$  and  $J_{\text{sc}}$ . For instance, adding 10 mM CDCA to **EDOT-Ph** increased  $V_{\text{oc}}$  from 0.59 to 0.75 V and  $J_{\text{sc}}$  from 2.7 to 8.2  $\text{mA cm}^{-2}$ . These changes support previous reports that CDCA improves the balance of electron injection over recombination by shifting the position of the conduction band.<sup>38,39</sup> Also, because these dyes are small, unbranched and highly conjugated they may aggregate on the  $\text{TiO}_2$  surface which is detrimental to device performance. However, CDCA suppresses dye aggregation<sup>38–40</sup> which also improves device efficiencies.

For the best dye (**EDOT-Ph**), the further addition of 0.05 mM LiI improves  $\eta$  very slightly (4.0% *vs.* 3.9%). Underlying this are larger changes; *e.g.*  $J_{\text{sc}}$  increases from 8.2 to 9.8  $\text{mA cm}^{-2}$  whilst  $V_{\text{oc}}$  drops from 0.75 to 0.62 V. This reflects  $\text{Li}^+$  ions adsorbing on the  $\text{TiO}_2$  surface and shifting the conduction band as reported previously.<sup>41,42</sup> This improves electron injection but at the expense of voltage. Adding LiI has little effect on **EDOT-Th** devices but it does improve the performance of **EDOT-Fu** from 2.4 to 2.9% for similar reasons to **EDOT-Ph** but the  $V_{\text{oc}}$  drops less. In addition, the dark currents for the three EDOT dyes are all very similar (Fig. S4). This suggests that the dyes exhibit similar rates of recombination at the electrolyte–electrode

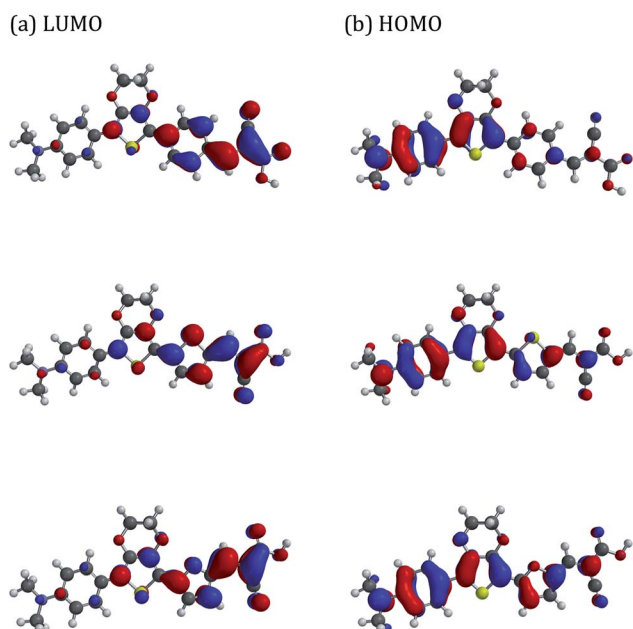


Fig. 2 HOMO/LUMO maps of dyes predicted by DFT calculations. Top: **EDOT-Ph**, middle: **EDOT-Fu**, bottom: **EDOT-Th**.

Table 2 *I*–*V* data for EDOT dyes

Dye	LiI	CDCA	FF	$V_{\text{oc}}$ (V)	$J_{\text{sc}}$ ( $\text{mA cm}^{-2}$ )	$\eta$ (%)
<b>EDOT-Ph</b>	N	N	0.62	0.59	2.70	1.0
	N	Y	0.64	0.75	8.20	3.9
	Y	Y	0.62	0.62	9.80	4.0
<b>EDOT-Fu</b>	N	N	0.63	0.61	3.10	1.2
	N	Y	0.63	0.70	5.40	2.4
	Y	Y	0.62	0.61	7.70	2.9
<b>EDOT-Th</b>	N	N	0.34	0.54	2.50	0.5
	N	Y	0.62	0.70	6.50	2.7
	Y	Y	0.67	0.70	6.00	2.8
<b>N719</b>			0.67	0.73	12.30	6.0
<b>D205</b>			0.68	0.74	11.40	5.8



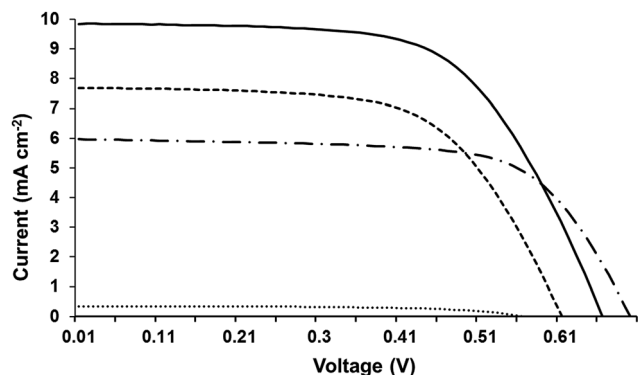


Fig. 3  $I$ - $V$  curves for EDOT-Ph (solid line), EDOT-Fu (dashed line), EDOT-Th (dot-dash line) and dark current for EDOT-Ph (dotted line).

interface and hence infers that a different factor is influencing device performance for these dyes.

Despite the 3 dyes all exhibiting the similar structural motifs and *N,N*-dimethylaniline donor groups, **EDOT-Ph** gives the best performance. The addition of a hetero-atom (O or S) into the linker unit does not increase device efficiency with the highest  $J_{sc}$  (Table 2) and EQE (Fig. 4) recorded for the **EDOT-Ph** dye containing a phenyl linker group. Interestingly, the EQE data show spectral response from 350 to 600 nm for all 3 dyes indicating that the heteroatoms in **EDOT-Fu** and **EDOT-Th** do not directly influence the  $\lambda$  range of light harvesting. However, UV-visible data for the dyes adsorbed onto  $\text{TiO}_2$  films (Fig. S2d†) show that light absorption is linked to  $J_{sc}$  with the highest  $J_{sc}$  and absorbance for EDOT-Ph ( $9.80 \text{ mA cm}^{-2}$  and  $\text{Abs} = 2.86$ ) and the lowest  $J_{sc}$  and absorbance for EDOT-Th ( $6.00 \text{ mA cm}^{-2}$  and  $\text{Abs} = 1.58$ ). Therefore, light harvesting appears to be strongly correlated with dye loading.

### Electronic spectra of adsorbed dyes

UV-vis spectra (Fig. S2†) have been studied for the three EDOT dyes adsorbed from 0.5 mM THF solutions by 18 h passive

dyeing onto transparent, mesoporous  $\text{TiO}_2$  films produced by doctor blading 1 layer of DSL18NR-T colloid onto glass (film thickness *ca.* 7  $\mu\text{m}$ ). The absorbance data have been measured in the presence of varying concentrations of CDCA to study the influence of this additive on dye aggregation. The data show  $\lambda_{\text{max}}$  at *ca.* 440 nm for **EDOT-Ph** and **EDOT-Fu** but a red-shift for **EDOT-Th** to *ca.* 550 nm reflecting the addition of the sulphur heteroatom in the thiophene moiety of this dye. In addition, each of the 3 dyes exhibit similar trends of narrowing of the main absorption peak as the CDCA concentration increases. Taking the **EDOT-Fu** data as an example (Fig. S2†), this shows the broadest absorption in the absence of any CDCA; *e.g.*  $\text{Abs} > 0.9$  stretches from 360 to 550 nm. However, even the addition of 1 mM CDCA results in a drastic narrowing of the absorption peak ( $\text{Abs} > 0.9$  is 400 to 505 nm) whilst at  $\geq 15$  mM CDCA  $\text{Abs} > 0.9$  narrows to 420 to 470 nm. This effect is ascribed to dye aggregation which is at its greatest when there are only EDOT molecules present because the intermolecular dye interactions are also at their strongest which causes a significant broadening of the main absorption peak. However, when 1 mM CDCA is added to the dye solution, there are 2 CDCA molecules for every EDOT dye and this significantly reduces the inter-dye interactions. This leads to fewer aggregated dye molecules when they adsorb on the  $\text{TiO}_2$  surface. This is important because dye aggregation has previously been shown to reduce device efficiency (*e.g.* for half-squarylium<sup>12,13</sup> and squaraine dyes).<sup>38,43</sup> So, at first sight, the broader light absorption without CDCA might seem beneficial but the poorer electron injection and dye regeneration combined with increased inter-dye recombination between aggregated dye molecules are believed to outweigh the enhanced light harvesting. When comparing the different dyes, CDCA has the greatest influence on **EDOT-Fu** which suggests greater aggregation for this dye. This may be due to the oxygen of the furan moiety causing greater polarisation within this dye and add to the  $\pi$ - $\pi$  stacking interactions which are likely between these small, unbranched but highly conjugated dye molecules.

## Experimental

### General synthetic, analytical and computational procedures

All synthetic transformations were performed under a nitrogen atmosphere. Solvents were purified using a PureSolv solvent purifier system. Melting points are uncorrected.  $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectroscopy were recorded on a Bruker AV III (500 MHz) spectrometer, operating at 500 MHz and 125 MHz, respectively. Chemical shifts are reported in ppm and are relative to TMS. UV-vis spectra were recorded using a Perkin Elmer Lambda 25 spectrophotometer. Square wave voltammetry was recorded at room temperature under nitrogen on a CH-Instruments 440A potentiostat using a three-electrode cell with a Pt working electrode, a Pt wire counter electrode and an Ag wire pseudo-reference electrode.  $\text{Bu}_4\text{NPF}_6$  (0.1 M) was used as the supporting electrolyte in DMF. All voltages are referenced to  $\text{Fc}/\text{Fc}^+$  redox couple and are adjusted to 0 V. DFT calculations were performed using Spartan '14 (64 bit).<sup>44</sup> The structures were first optimized semi-empirically (AM1) and then re-optimized

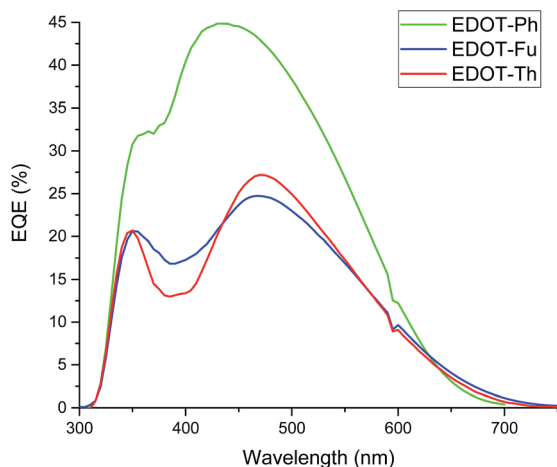


Fig. 4 EQE data for EDOT-Ph (green line), EDOT-Fu (red line) and EDOT-Th (blue line) each with 10 mM CDCA and 0.05 mM Lil ( $1 \text{ cm}^2$  devices).





using DFT (B3LYP/6-31G\*). The resulting structures were local minima, as none of the vibrational frequencies generated imaginary frequencies.

## Synthesis

**General procedure for synthesis of compounds 4a–c.** 4-Bromo-*N,N*-dimethylaniline (0.32 g, 1.6 mmol), EDOT (1) (0.33 mL, 3.1 mmol), cesium carbonate (1.32 g, 4.0 mmol) palladium(II) acetate (18.3 mg, 0.08 mmol), tri-*tert*-butylphosphine tetrafluoroborate (70 mg, 0.24 mmol), pivalic acid (65 mg, 0.6 mmol) were introduced to an oven dried 2-neck flask under an argon atmosphere. Anhydrous DMF (5 mL) was added *via* a syringe and the reaction mixture was allowed to stir at 100 °C for 18 h. After this time aldehyde 3a–c (0.60 g, 3.2 mmol), palladium(II) acetate (18.3 mg, 0.08 mmol), tri-*tert*-butylphosphine tetrafluoroborate (70 mg, 0.24 mmol) were added to the reaction mixture and left to stir at 100 °C for 24 h. The mixture was cooled to room temperature and passed through a pad of Celite 545 and washed with DCM. The filtrate was concentrated under reduced pressure. Purification of 4a–c was achieved by column chromatography on silica gel (petroleum ether : DCM; 1 : 2).

Compound 4a was prepared using aldehyde 3a and was isolated as a red solid (341.6 mg, 58%). Mp 155–156 °C; <sup>1</sup>H NMR (CDCl<sub>3</sub>, 500 MHz, ppm) δ<sub>H</sub> 9.95 (1H, s, CHO), 7.88 (2H, d, *J* 8.5, CH-Ar), 7.84 (2H, d, *J* 8.5, CH-Ar), 7.65 (2H, d, *J* 9.0, CH-Ar), 6.74 (2H, d, *J* 9.0, CH-Ar), 4.39–4.41 (2H, m, CH<sub>2</sub>), 4.34–4.36 (2H, m, CH<sub>2</sub>), 3.00 (6H, s, 2 × CH<sub>3</sub>); <sup>13</sup>C NMR (CDCl<sub>3</sub>, 125 MHz, ppm) δ<sub>C</sub> 191.5 (CHO), 149.7 (C–NH<sub>2</sub>), 141.0 (C–OCH<sub>2</sub>), 139.6 (C–OCH<sub>2</sub>), 137.1 (C–CHO), 133.4 (C), 130.2 (CH-Ar), 127.4 (CH-Ar), 125.5 (CH-Ar), 120.7 (C), 119.6 (C), 112.3 (CH-Ar), 111.7 (C), 64.9 (CH<sub>2</sub>), 64.4 (CH<sub>2</sub>), 40.5 (2 × CH<sub>3</sub>); HRMS *m/z* (ESI<sup>+</sup>) [M]<sup>+</sup> 365.1065 (requires 365.1080 for C<sub>21</sub>H<sub>19</sub>NO<sub>3</sub>S).

Compound 4b was prepared using aldehyde 3b and was isolated as a red solid (181 mg, 32%). Mp 146–147 °C; <sup>1</sup>H NMR (CDCl<sub>3</sub>, 500 MHz, ppm) δ<sub>H</sub> 9.55 (1H, s, CHO), 7.63 (2H, d, *J* 9.0, 2 × CH-Ar), 7.31 (1H, d, *J* 3.8, CH-Fu), 6.74 (3H, m, 2 × CH-Ar and CH-Fu), 4.42 (2H, dt, *J* 3.9, 2.0, CH<sub>2</sub>), 4.36 (2H, dt, *J* 3.9, 2.0, CH<sub>2</sub>), 3.01 (6H, s, 2 × CH<sub>3</sub>); <sup>13</sup>C NMR (CDCl<sub>3</sub>, 125 MHz, ppm) δ<sub>C</sub> 176.4 (CHO), 154.4 (C), 150.5 (C), 150.2 (C), 141.9 (C), 136.8 (C), 127.8 (CH-Ar), 127.8 (CH-Fu), 121.8 (C), 120.8 (C), 112.7 (CH-Ar), 108.3 (CH-Fu), 102.6 (C), 65.5 (CH<sub>2</sub>), 64.8 (CH<sub>2</sub>), 40.8 (2 × CH<sub>3</sub>); HRMS *m/z* (ESI<sup>+</sup>) [M + Na]<sup>+</sup> 378.0754 (requires 378.0770 for C<sub>19</sub>H<sub>17</sub>NNaO<sub>4</sub>S).

Compound 4c was prepared using aldehyde 3c and was isolated as a red solid (215 mg, 36%). Mp 140–141 °C; <sup>1</sup>H NMR (CDCl<sub>3</sub>, 500 MHz, ppm) δ<sub>H</sub> 9.84 (1H, s, CHO), 7.65 (1H, d, *J* 4.1, CH-Th), 7.62 (2H, d, *J* 9.1, CH-Ar), 7.23 (1H, d, *J* 4.1, CH-Th), 6.73 (2H, d, *J* 9.1, CH-Ar), 4.42–4.44 (2H, m, CH<sub>2</sub>), 4.34–4.37 (2H, m, CH<sub>2</sub>), 3.00 (6H, s, 2 × CH<sub>3</sub>); <sup>13</sup>C NMR (CDCl<sub>3</sub>, 125 MHz, ppm) δ<sub>C</sub> 182.7 (CHO), 150.2 (C), 145.9 (C), 141.4 (C), 140.4 (C), 137.4 (CH-Th), 137.0 (C), 127.8 (CH-Ar), 122.5 (CH-Th), 120.7 (C), 120.3 (C), 112.7 (CH-Ar), 107.4 (C), 65.6 (CH<sub>2</sub>), 64.9 (CH<sub>2</sub>), 40.8 (CH<sub>3</sub>); HRMS *m/z* (ESI<sup>+</sup>) [M + Na]<sup>+</sup> 394.0559 (requires 394.0548 for C<sub>19</sub>H<sub>17</sub>NNaO<sub>3</sub>S<sub>2</sub>).

**General procedure for the synthesis of EDOT-Ph/Fu/Th.** Compounds 4a–c (1.37 mmol) and cyanoacetic acid (0.93 g,

11.0 mmol) were dissolved in a 1 : 1 mixture of toluene and acetonitrile (10 mL + 10 mL) in the presence of piperidine (1.2 mL, 21.0 mmol), and then heated to 85–90 °C for 6 h. The mixture was then cooled to room temperature, and solvent was removed under reduced pressure. The residue was purified by column chromatography on silica gel (DCM : MeOH; 4 : 1).

**EDOT-Ph** was isolated as a red powder (520 mg, 88%). Mp > 250 °C; <sup>1</sup>H NMR (DMSO-*d*<sub>6</sub>, 400 MHz, ppm) δ<sub>H</sub> 7.93 (1H, s, CH=C), 7.90 (2H, d, *J* 8.5, 2 × CH-Ar), 7.77 (2H, d, *J* 8.5, 2 × CH-Ar), 7.56 (2H, d, *J* 9.0, 2 × CH-Ar), 6.77 (2H, d, *J* 9.0, 2 × CH-Ar), 4.43 (2H, dd, *J* 5.8, 2.1, CH<sub>2</sub>), 4.38 (2H, dd, *J* 5.8, 2.1, CH<sub>2</sub>), 2.94 (6H, s, 2 × CH<sub>3</sub>); HRMS *m/z* (ESI<sup>−</sup>) [M − H]<sup>−</sup> 431.1061 (requires 431.1071 for C<sub>24</sub>H<sub>19</sub>N<sub>2</sub>O<sub>4</sub>S).

**EDOT-Fu** was isolated as a red powder (532 mg, 90%). Mp > 250 °C; <sup>1</sup>H NMR (DMSO-*d*<sub>6</sub>, 500 MHz, ppm) δ<sub>H</sub> 7.69 (1H, s, CH=C), 7.54 (2H, d, *J* 9.1, 2 × CH-Ar), 7.25 (1H, d, *J* 3.7, CH-Fu), 6.78 (2H, d, *J* 9.1, 2 × CH-Ar), 6.76 (1H, d, *J* 3.7, CH-Fu), 4.43–4.45 (2H, m, CH<sub>2</sub>), 4.38–4.40 (2H, m, CH<sub>2</sub>), 2.95 (6H, s, 2 × CH<sub>3</sub>); HRMS *m/z* (ESI<sup>−</sup>) [M − H]<sup>−</sup> 421.0875 (requires 421.0864 for C<sub>22</sub>H<sub>17</sub>N<sub>2</sub>O<sub>5</sub>S).

**EDOT-Th** as isolated as a red powder (490 mg, 83%). Mp 240 °C (dec); <sup>1</sup>H NMR (DMSO-*d*<sub>6</sub>, 500 MHz, ppm) δ<sub>H</sub> 8.13 (1H, d, *J* 0.4, CH=C), 7.65 (1H, d, *J* 4.1, CH-Th), 7.56 (2H, d, *J* 9.1, 2 × CH-Ar), 7.28 (1H, d, *J* 4.1, CH-Th), 6.77 (2H, d, *J* 9.1, 2 × CH-Ar), 4.47 (2H, dt, *J* 3.8, 2.1, CH<sub>2</sub>), 4.40 (2H, dt, *J* 3.8, 2.1, CH<sub>2</sub>), 2.95 (6H, s, 2 × CH<sub>3</sub>); HRMS *m/z* (ESI<sup>−</sup>) [M − H]<sup>−</sup> 437.0626 (requires 437.0635 for C<sub>22</sub>H<sub>17</sub>N<sub>2</sub>O<sub>4</sub>S<sub>2</sub>).

**One-pot synthesis of EDOT-Ph.** EDOT (1) (0.30 mL, 2.81 mmol), 4-bromo-*N,N*-dimethylaniline (2) (0.68 g, 3.39 mmol), cesium carbonate (0.68 g, 2.08 mmol) palladium(II) acetate (32 mg, 0.14 mmol), tricyclohexylphosphine tetrafluoroborate (0.21 g, 0.57 mmol) and pivalic acid (48 μL, 0.41 mmol) were dissolved in anhydrous DMF (2.5 mL) and the reaction mixture was allowed to stir at 100 °C for 24 h. *p*-Bromobenzylidenecyanoacetic acid (5) (0.86 g, 3.41 mmol), cesium carbonate (0.68 g, 2.08 mmol) palladium(II) acetate (32 mg, 0.14 mmol), tricyclohexylphosphine tetrafluoroborate (0.21 g, 0.57 mmol), pivalic acid (48 μL, 0.41 mmol) dissolved in anhydrous DMF (2.5 mL) were then added and the mixture was stirred for 24 h at 100 °C. The mixture was cooled to room temperature and passed through a layer of Celite 545 using DCM. The filtrate was washed with water (50 mL) and the aqueous layer was extracted with DCM (2 × 30 mL). The combined organic layers were washed with brine, dried over MgSO<sub>4</sub>, filtered and concentrated under reduced pressure. Purification by column chromatography on silica gel (DCM : MeOH; 5 : 1) afforded **EDOT-Ph** as a red solid (0.31 g, 26%).

## Conclusions

Our one-pot, three component C–H activation protocols have allowed the rapid synthesis of precursors for three EDOT-containing dyes. Moreover, we have extended this methodology to produce **EDOT-Ph** in a respectable yield following facile purification using column chromatography. In the longer term, our protocols could be further extended to provide



a combinatorial-like approach to allow the rapid screening of the structure–property relationships of diverse libraries of dyes.<sup>35,45</sup>

Interestingly, a comparison of device performance data of the three dyes shows that the *N,N*-dimethylaniline moiety produces devices with comparable efficiency to analogous sensitizers containing a *N,N*-dimethylaniline donor unit attached to the EDOT unit *via* an acetylene bridge or dyes containing more complicated indoline or triphenylamine donor units reported in the literature.<sup>26,28,37</sup> A substantial improvement in device performance (8.2%) was reported for a triphenylamine containing EDOT sensitizer but this is due to a modification of the acceptor unit in the dye.<sup>37</sup> Therefore, a comparison of the EDOT dyes reported in the literature with the EDOT dyes reported in this work would suggest that the *N,N*-dimethylaniline moiety is not a significantly poorer donor group than triphenylamine. However, we do observe unusually high device efficiency improvements on addition of CDCA. For example, optimised CDCA concentrations required for the sensitizers reported in this study are 10 mM by comparison to 1 mM and 6 mM for sensitizers reported in other studies<sup>38,39</sup> respectively. This suggests that the bulkier indoline and triphenylamine groups may reduce dye aggregation more effectively than the *N,N*-dimethylaniline unit. At the other end of the dye molecules, the phenyl-linked dye **EDOT-Ph** gives rise to better device performance than the heterocyclic variants despite the latter extending their light harvesting to longer wavelengths.

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